

**2410U10-1A**

**TUESDAY, 14 MAY 2024 – MORNING**

**CHEMISTRY – AS unit 1**

**DATA BOOKLET**

Avogadro constant	$N_A$	=	$6.02 \times 10^{23} \text{ mol}^{-1}$
molar gas constant	$R$	=	$8.31 \text{ J mol}^{-1} \text{ K}^{-1}$
molar gas volume at 273 K and 1 atm	$V_m$	=	$22.4 \text{ dm}^3 \text{ mol}^{-1}$
molar gas volume at 298 K and 1 atm	$V_m$	=	$24.5 \text{ dm}^3 \text{ mol}^{-1}$
Planck constant	$h$	=	$6.63 \times 10^{-34} \text{ Js}$
speed of light	$c$	=	$3.00 \times 10^8 \text{ m s}^{-1}$
density of water	$d$	=	$1.00 \text{ g cm}^{-3}$
specific heat capacity of water	$c$	=	$4.18 \text{ J g}^{-1} \text{ K}^{-1}$
ionic product of water at 298 K	$K_w$	=	$1.00 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6}$
fundamental electronic charge	$e$	=	$1.60 \times 10^{-19} \text{ C}$

temperature (K) = temperature ( $^{\circ}\text{C}$ ) + 273

$1 \text{ dm}^3 = 1000 \text{ cm}^3$

$1 \text{ m}^3 = 1000 \text{ dm}^3$

$1 \text{ tonne} = 1000 \text{ kg}$

$1 \text{ atm} = 1.01 \times 10^5 \text{ Pa}$

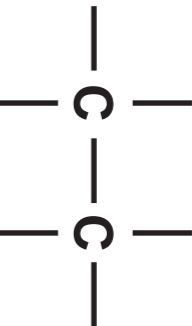
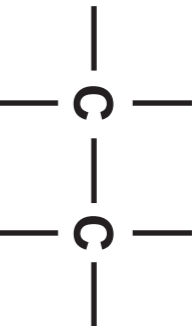
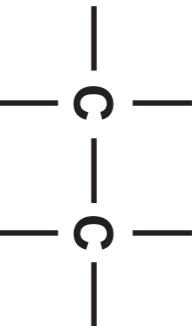
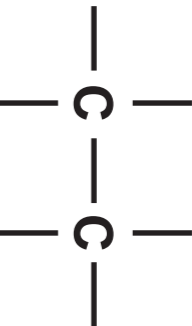

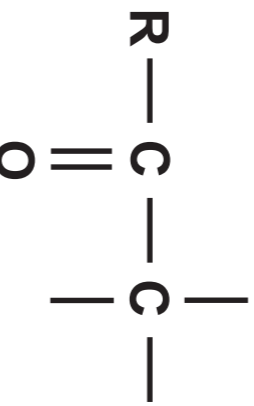
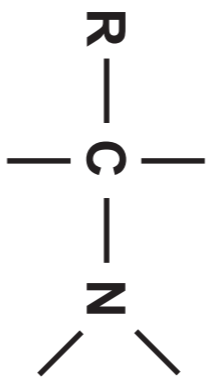
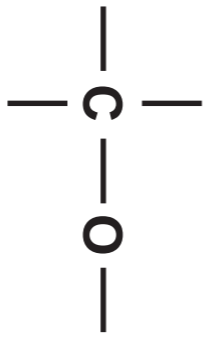





Multiple	Prefix	Symbol
$10^{-9}$	nano	n
$10^{-6}$	micro	$\mu$
$10^{-3}$	milli	m

Multiple	Prefix	Symbol
$10^3$	kilo	k
$10^6$	mega	M
$10^9$	giga	G





## INFRARED ABSORPTION VALUES

<b>BOND</b>	<b>WAVENUMBER / cm<sup>-1</sup></b>
<b>C—Br</b>	<b>500 to 600</b>
<b>C—Cl</b>	<b>650 to 800</b>
<b>C—O</b>	<b>1000 to 1300</b>
<b>C=C</b>	<b>1620 to 1670</b>
<b>C=O</b>	<b>1650 to 1750</b>
<b>C≡N</b>	<b>2100 to 2250</b>
<b>C—H</b>	<b>2800 to 3100</b>
<b>O—H (carboxylic acid)</b>	<b>2500 to 3200 (very broad)</b>
<b>O—H (alcohol / phenol)</b>	<b>3200 to 3550 (broad)</b>
<b>N—H</b>	<b>3300 to 3500</b>

**<sup>13</sup>C NMR CHEMICAL SHIFTS RELATIVE TO TMS = 0**

<b>TYPE OF CARBON</b>	<b>CHEMICAL SHIFT, <math>\delta</math> (ppm)</b>
	<b>5 to 40</b>
	<b>5 to 40</b>
	<b>5 to 40</b>
	<b>5 to 40</b>
	<b>10 to 70</b>
	<b>20 to 50</b>
	<b>25 to 60</b>
	<b>50 to 90</b>
	<b>90 to 150</b>
	<b>110 to 125</b>
	<b>110 to 160</b>
	<b>160 to 185</b>
	<b>190 to 220</b>

# <sup>1</sup>H NMR CHEMICAL SHIFTS RELATIVE TO TMS = 0

TYPE OF PROTON	CHEMICAL SHIFT, $\delta$ (ppm)
$-\text{CH}_3$	0.1 to 2.0
$\text{R}-\text{CH}_3$	0.9
$\text{R}-\text{CH}_2-\text{R}$	1.3
$\text{CH}_3-\text{C}\equiv\text{N}$	2.0
$\text{CH}_3-\text{C}(=\text{O})-$	2.0 to 2.5
$-\text{CH}_2-\text{C}(=\text{O})-$	2.0 to 3.0
 $\text{C}_6\text{H}_5-\text{CH}_3$	2.2 to 2.3
$\text{HC}-\text{Cl}$ or $\text{HC}-\text{Br}$	3.1 to 4.3
$\text{HC}-\text{O}$	3.3 to 4.3
$\text{R}-\text{OH}$	4.5 *
$-\text{C}=\text{CH}$	4.5 to 6.3
$-\text{C}=\text{CH}-\text{CO}$	5.8 to 6.5
 $\text{C}_6\text{H}_5-\text{CH}=\text{C}$	6.5 to 7.5
 $\text{C}_6\text{H}_5-\text{H}$	6.5 to 8.0
 $\text{C}_6\text{H}_5-\text{OH}$	7.0 *
$\text{R}-\text{C}(=\text{O})-\text{H}$	9.8 *
$\text{R}-\text{C}(=\text{O})-\text{OH}$	11.0 *

\*variable figure dependent on concentration and solvent

# THE PERIODIC TABLE OF ELEMENTS

## KEY

Ar	relative atomic mass
Sym	symbol
Z	atomic number

1		2		GROUP										3	4	5	6	7	0																	
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7 Li 3	9 Be 4											11 B 5	12 C 6	14 N 7	16 O 8	19 F 9	20 Ne 10																			
23 Na 11	24 Mg 12											27 Al 13	28 Si 14	31 P 15	32 S 16	35.5 Cl 17	40 Ar 18																			
39 K 19	40 Ca 20	45 Sc 21	48 Ti 22	51 V 23	52 Cr 24	55 Mn 25	56 Fe 26	59 Co 27	59 Ni 28	63.5 Cu 29	65 Zn 30	70 Ga 31	73 Ge 32	75 As 33	79 Se 34	80 Br 35	84 Kr 36																			
86 Rb 37	88 Sr 38	89 Y 39	91 Zr 40	93 Nb 41	96 Mo 42	99 Tc 43	101 Ru 44	103 Rh 45	106 Pd 46	108 Ag 47	112 Cd 48	115 In 49	119 Sn 50	122 Sb 51	128 Te 52	127 I 53	131 Xe 54																			
133 Cs 55	137 Ba 56	139 La 57	179 Hf 72	181 Ta 73	184 W 74	186 Re 75	190 Os 76	192 Ir 77	195 Pt 78	197 Au 79	201 Hg 80	204 Tl 81	207 Pb 82	209 Bi 83	210 Po 84	210 At 85	222 Rn 86																			
223 Fr 87	226 Ra 88	227 Ac 89																																		

**THE PERIODIC TABLE – KEY**  
**ATOMIC NUMBER – SYMBOL – NAME**

1	H - Hydrogen	38	Sr - Strontium	75	Re - Rhenium
2	He - Helium	39	Y - Yttrium	76	Os - Osmium
3	Li - Lithium	40	Zr - Zirconium	77	Ir - Iridium
4	Be - Beryllium	41	Nb - Niobium	78	Pt - Platinum
5	B - Boron	42	Mo - Molybdenum	79	Au - Gold
6	C - Carbon	43	Tc - Technetium	80	Hg - Mercury
7	N - Nitrogen	44	Ru - Ruthenium	81	Tl - Thallium
8	O - Oxygen	45	Rh - Rhodium	82	Pb - Lead
9	F - Fluorine	46	Pd - Palladium	83	Bi - Bismuth
10	Ne - Neon	47	Ag - Silver	84	Po - Polonium
11	Na - Sodium	48	Cd - Cadmium	85	At - Astatine
12	Mg - Magnesium	49	In - Indium	86	Rn - Radon
13	Al - Aluminium	50	Sn - Tin	87	Fr - Francium
14	Si - Silicon	51	Sb - Antimony	88	Ra - Radium
15	P - Phosphorus	52	Te - Tellurium	89	Ac - Actinium
16	S - Sulfur	53	I - Iodine	90	Th - Thorium
17	Cl - Chlorine	54	Xe - Xenon	91	Pa - Protactinium
18	Ar - Argon	55	Cs - Caesium	92	U - Uranium
19	K - Potassium	56	Ba - Barium	93	Np - Neptunium
20	Ca - Calcium	57	La - Lanthanum	94	Pu - Plutonium
21	Sc - Scandium	58	Ce - Cerium	95	Am - Americium
22	Ti - Titanium	59	Pr - Praseodymium	96	Cm - Curium
23	V - Vanadium	60	Nd - Neodymium	97	Bk - Berkelium
24	Cr - Chromium	61	Pm - Promethium	98	Cf - Californium
25	Mn - Manganese	62	Sm - Samarium	99	Es - Einsteinium
26	Fe - Iron	63	Eu - Europium	100	Fm - Fermium
27	Co - Cobalt	64	Gd - Gadolinium	101	Md - Mendeleevium
28	Ni - Nickel	65	Tb - Terbium	102	No - Nobelium
29	Cu - Copper	66	Dy - Dysprosium	103	Lr - Lawrencium
30	Zn - Zinc	67	Ho - Holmium		
31	Ga - Gallium	68	Er - Erbium		
32	Ge - Germanium	69	Tm - Thulium		
33	As - Arsenic	70	Yb - Ytterbium		
34	Se - Selenium	71	Lu - Lutetium		
35	Br - Bromine	72	Hf - Hafnium		
36	Kr - Krypton	73	Ta - Tantalum		
37	Rb - Rubidium	74	W - Tungsten		